

Molarity Pogil Answers

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Many companies, including pharmaceutical companies, keep track of the concentration of a solution by measuring its molarity – a ratio of the number of solute particles to the volume of the solution. In this activity you will learn about molarity and how to represent concentration quantitatively. Model 1 – Lemonade Mixtures*

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of the new solution? (Answer: 9-801 HN03) = ML (0.8M -N 3-30.00 ml, of water is added to 20.00 ml, of 0.0500 M sodium carbonate. What is the molarity of sodium carbonate in the new solution? (Answer: 0.0200 M) (20.00 m C (350. 00 casF) 4. What volume of 0.100 M NaCl is needed to make 5.00 ml, of 0.200 M solution? (Answer: 0.01 (0.100 M) V)

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of moles of solute (mol) divided by the total volume of the solution in liters (L). Recall that that volume of the solution is the sum of the volume of solvent and the volume of solute. $\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} = M = \text{mol/L}$ Revised 2017-2018 LCA

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